

Antibacterial evaluation of Some Synthesis Complexes of Azo-dye Ligands.

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Abstract

In present work, two ligands, azo Schiff-base (AZD1 and AZD2), were synthesized by two steps. Firstly, Schiff-bases were prepared by condensation of Salicylaldehyde with anthranilic acid, followed by diazotization of aromatic amines and their subsequent coupling reaction with Schiff-bases. Consequently, 8 complexes of the type $[M(AZD1)Cl_2]$ and $[M(AZD2)Cl_2]$, where $M = Cu(II), Ni(II), Co(II), Mn(II)$; were synthesized and characterized with the help of molar conductance, magnetic susceptibility measurements, FT-IR, micro elemental analysis, 1H NMR, and Uv-visible spectroscopy. The spectral data revealed that the ligands act as bidentate ligand, the hydroxyl of carboxylic group and the azo group in AZD1 and hydroxyl of phenyl group and the azomethine group in AZD2 ligand are co-ordination to the central metal atoms. Lastly, all the compounds were screened for their antibacterial activity, using well-diffusion method. In conclusion, the ligands and some of the complexes exhibited appreciable antibacterial activity against *Escherichia coli*, *Klebsiella pneumonia* and *Staphylococcus haemolyticus* and *Staphylococcus aureus*.

KeyWords: Azo Schiff base, Azo Schiff base complexes, antibacterial activity and spectral studies.

1.0 Introduction

Azo Schiff base complexes contain both azo and azomethine groups. The azo group has an excellent, important application in coordination chemistry^(1, 2). It is a highly sensitive organic reagent, suitable for the spectrophotometric determination of ions of transition elements in trace concentrations⁽³⁾. They also have a wide application as metallochromic indicators for colorometric titration⁽⁴⁾.

These dyes are used in electrophotographic or sensor applications, lasers, electro-optical devices and ink-jet printers ⁽⁵⁾. The coordination complexes of transition metals with azo ligands are of current attraction due to the interesting physical, chemical, photophysical and photochemical, catalytic and different material properties ⁽⁶⁾.

Azo compounds or dyes characterized by the presence of the azo moiety ($-N=N-$) in their structure, conjugated with two, distinct or identical, mono- or polycyclic aromatic or hetero aromatic systems ⁽⁷⁾. Previous studies have shown the antimicrobial activity of metal complexes against various pathogens including bacteria ⁽⁸⁾. However, the antimicrobial activity of any synthetic metal complexes might be different and therefore, it should be revealed.

In the current study azo dyes were synthesized from reaction of (2-amino-benzoic acid) with salicylaldehyde and uses the (aniline or 4-nitroaniline) as coupling agent scheme (1). Then they were reacted with aniline and nitroaniline to prepare Schiff bases as ligands. In addition, complexes of this ligand with metal ions of Cu(II), Ni(II), Co(II) and Mn(II) have also been synthesized and characterized spectroscopically. Finally, the antimicrobial activity of these synthesized complexes was determined using disc diffusion assay.

2.0 Experimental

2.1 Instrumentation

The UV-visible spectrophotometer was carried out in ethanol and water mixture as solvent by using the T50 (Shimadzu) spectrophotometer, FTIR spectrophotometer were taken on a (Shimadzu) infrared spectrophotometer in the $(4000-400) \text{ cm}^{-1}$ range. Melting points were measured using Stuart Digital Melting point apparatus. Elemental analysis was carried out on Perkin Elmer-2400 CHNS analyzer. The conductivities of the complexes were measured in ethanol and water mixture as solvent by using Conductometer type (Cond. 720) conductivity meter. In addition, magnetic measurements were recorded on an auto Magnetic Susceptibility balance Guoy method. Finally, Antibacterial evaluations were measured by using well diffusion methods against four types of bacteria.

2.2 Materials and Reagents

All the chemicals used were of reagent grade and procured from Sigma-Aldrich and Fluka. Metal salts (B.D.H) were commercially available as pure samples and solvents were used as received.

2.3 Preparation of the Ligands

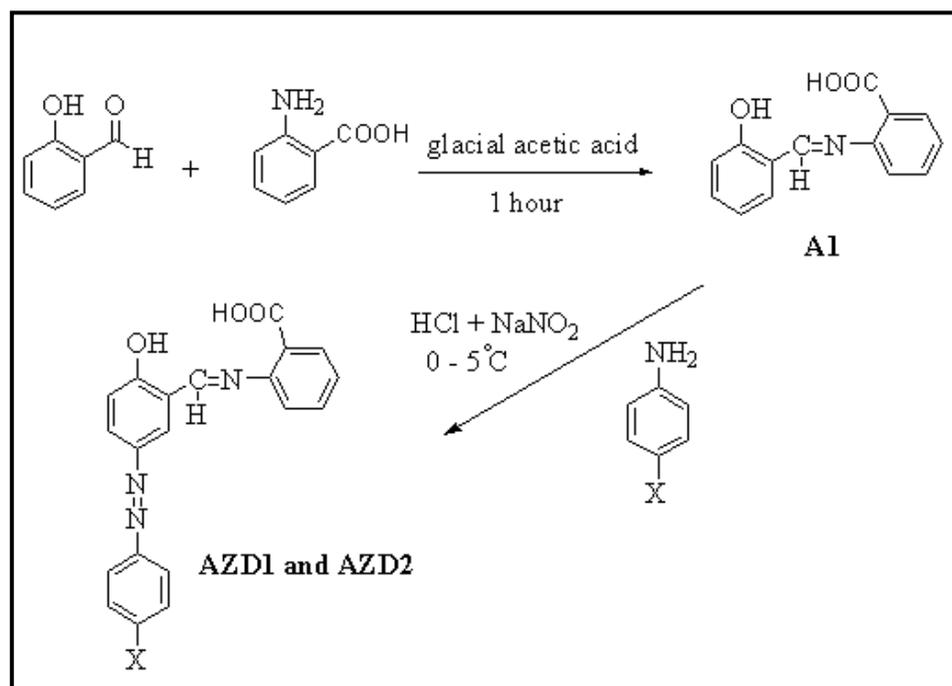
The ligands were prepared according to the procedure ⁽⁹⁾.

2.3.1 Preparation of Schiff bases (starting materials) (A1)

A mixture of (2-amino benzoic acid) (0.05 mol) and Salicylaldehyde (3.425 gm, 0.05 mol) in 96% ethanol (50 mL) with few drops of glacial acetic acid was stirred for 1 hour. The precipitate was filtrated, washed several times with ethanol and dried to give Salicylidene 2- aminobenzoic acid (A1). The reaction is shown in Scheme 1.

2.3.2 Synthesis of Azo Schiff base ligands (dyes) (AZD1 and AZD2)

(0.01mol) of (aniline or 4- nitroaniline) was gently heated in (1.7 mL) Cone. HCl and then (10 mL) of distilled water was added. The solution was then cooled to (0-5°C) in an ice-bath and maintained at this temperature. Sodium nitrite (0.69 gm) solution in water (5 mL) was then added dropwise. The diazonium solution was added portion wise to the coupling component solution prepared by mixing of (A1) (0.01mol) in water with sodium hydroxide (1.5gm) dissolve in (100 mL) of water. The mixture was stirred for 30 min. The precipitated crude dyes were formed, collected by filtration and washed with water and ethanol ⁽¹⁰⁾. The reaction is shown in Scheme (1). The physical properties of compounds were given in Table 1.



Scheme 1: Preparation of Ligands AZD1 and AZD2.

Table 1: Same physical properties of starting materials

No.	Compound	Attached group	Chem. Formula	Mol. Weight	Yield %	Colour
1	A1		C ₁₄ H ₁₁ NO ₃	241.242	75.5	Bright - Yellow
2	AZD1	X: 4-H	C ₂₀ H ₁₅ N ₃ O ₃	345.351	85	Brown
3	AZD2	X: 4-NO ₂	C ₂₀ H ₁₄ N ₄ O ₅	390.349	74	Reddish- brown

2.3.3. Synthesis of Metal Complexes (General procedure):

An ethanolic solution of the ligands (AZD1 and AZD2) (1 mmol) (25 mL) was added gradually to the ethanolic solution (10 mL) (1 mmol) of CuCl₂.6H₂O, NiCl₂ CoCl₂ and MnCl₂.4H₂O respectively. i.e. (1:1) mole ratio for (metal to ligand). Then added 3 drops of triethylamine (TEA). The mixture was left under reflux with continuous stirring about (3-4) hours or about 1 day where upon the solid complexes precipitated. The resulting solid was washed with ethanol followed by diethyl ether and dried in a vacuum over anhydrous CaCl₂.

2.4 Antibacterial Activity:

The bacterial strains were isolated by culturing from patients in Kalar General Hospital. The Mueller Hinton agar plates were cultured with 50µl of each type of bacteria (*E.coli* and *Staphylococcus aureus*) with OD₆₀₀ = 0.2 and were provided with (5) cm holes. The holes were loaded with (50) µl of all metal complexes and ligands, and the plates were incubated at 37 °C for (24) hours. Meropenem antibiotic was used as positive control and the solvent as a negative control. The diameter of the wells with inhibition zones were measured using ruler. The data were arranged in table (5).

2.0 Results and Discussion:

The structures of the ligand and the complexes were established from their Uv-Visible, FT- IR, elemental analyses, and magnetic susceptibility measurements. The complexes are intensely colored stable solids; the metal complexes exhibit (1: 1) metal to ligand stoichiometry, table (2). Molar conductivities of all synthesized complexes are listed in table (2), which compared with known molar conductivities^(11, 12). The low value of conductivities indicates the non-electrolyte nature of complexes.

Table (2): Color, melting point, elemental analysis and conductivity data for the synthesized azo- Schiff base metal complexes.

No.	Compound	Color (M.P).($^{\circ}\text{C}$)	Found (Calculated)%			m Λ ($\Omega^{-1} \cdot \text{cm}^2 \cdot \text{mol}^{-1}$)
			C %	H %	N %	
	Ligand(AZD1)	orange (182)	62.01 (61.38)	3.98 (3.86)	13.76 (14.32)	
1	[Cu(AZD1)Cl ₂]	Deep-Brown (324)	46.04 (45.77)	3.02 (2.79)	9.18 (10.65)	1
2	[Ni(AZD1)Cl ₂]	Orange-yellow (234)	45.94 (46.20)	3.00 (2.71)	11.02 (10.78)	1,5
3	[Co(AZD1)Cl ₂]	Deep-Brown (221)	45.33 (46.1)	2.93 (2.70)	9.54 (10.77)	1.9
4	[Mn(AZD1)Cl ₂]	Brown (156)	45.98 (46.54)	2.65 (3.97)	11.03 (10.85)	7.2
	Ligand (AZD2)	Brown (106)	60.95 (61.38)	3.27 (3.86)	15.01 (14.32)	
5	[Cu(AZD2)Cl ₂]	Orange- Brown (243)	50.87 (51.05)	3.59 (3.15)	9.18 (8.7)	10.9
6	[Ni(AZD2)Cl ₂]	Orange- Brown (122)	50.27 (50.12)	3.05 (3.18)	8.99 (8.10)	1.3
7	[Co(AZD2)Cl ₂]	Orange (154)	49.00 (50.92)	2.93 (3.18)	9.03 (8.10)	1.8
8	[Mn(AZD2)Cl ₂]	Brown (124)	46.56 (50.9)	3.94 (3.20)	9.4 (8.92)	1.2

3.1 Electronic spectra and magnetic moment:

The absorption bands at 45455 and 31250 cm^{-1} attributed to ($\pi \rightarrow \pi^*$) and ($n \rightarrow \pi^*$) transition respectively originating from the azo Schiff base ligand and carboxylic group are observed in the spectrum of the free ligand.⁽¹¹⁾ The data arranged in table (3)

The observed magnetic moment of the Cu(II) complex (1 and 4) is 1.94 and 1.95 B.M. respectively, which confirms the square planar structure of these complex. Electronic spectrum of the Cu(II) complex displays bands in range (46511- 23809) cm^{-1} . May be assigned to the charge transfer transitions⁽¹²⁾. The electronic spectrum of the orange yellow color complexes of Ni(II) ion (i.e. 2 and 6), has a strong band in range (22,222-45455) cm^{-1} located inside the experimental value at (15,000-25,000 cm^{-1}) of square planar geometry. This band ascribe to (${}^1A_{1g} \rightarrow {}^1A_{2g}$) transition. The magnetic measurements come also to confirm the square planar structure with a diamagnetic value⁽¹³⁾.

For Co(II) complexes (3) exhibit tow band at (22222 and 45455) assigned to charge transfer transitions of tetrahedral geometry. While Co(II) complexes (7), two bands at (45454 and 25000)

cm^{-1} assigned to charge transfer and ${}^4A_2(F) \rightarrow {}^4T_1(F)$ transitions, respectively. The effective magnetic moment for a tetrahedral Co(II) species is 4.06 B.M., while the calculated value is 3.93 and 3.89 B.M. for (3 and 7) respectively, this give a close matching data between experimental and theoretical data⁽¹⁴⁾.

The spectra of the complexes of Mn(II) shows an intense bands in range (23800 - 47615) cm^{-1} of tetrahedral geometry may be assigned to charge transfer transitions^(13, 15). The data of magnetic susceptibility confirm with the 5 unpaired electrons in d^5 system for metal complexes.

Table (3): The electronic and magnetic properties data for azo-Schiff base ligands and its metal complexes.

No.	Compounds	Band absorption		Assignment	μ_{eff} (B.M) Geometry
		cm^{-1}	nm		
	Ligand (AZD1)	<u>31250</u>	<u>320</u>	$\pi \rightarrow \pi^*$	
		<u>45450</u>	<u>235</u>	$n \rightarrow \pi^*$	
1	[Cu(AZD1)Cl ₂]	<u>46511</u>	<u>215</u>	CT	1.94 S.P.
		<u>40000</u>	<u>250</u>	CT	
		<u>23809</u>	<u>420</u>	CT	
2	[Ni(AZD1)Cl ₂]	<u>45455</u>	<u>220</u>	CT	0 S.P.
		<u>22222</u>	<u>450</u>	${}^1A_{1g} \rightarrow {}^1A_{2g}$	
3	[Co(AZD1)Cl ₂]	<u>45455</u>	<u>220</u>	CT	3.93 T.h.
		<u>22222</u>	<u>435</u>	CT	
4	[Mn(AZD1)Cl ₂]	<u>47619</u>	<u>210</u>	CT	4.95 T.h.
		<u>40000</u>	<u>250</u>	CT	
		<u>23809</u>	<u>420</u>	CT	
	Ligand (AZD2)	<u>45455</u>	<u>220</u>	$\pi \rightarrow \pi^*$	
		<u>23809</u>	<u>420</u>	$n \rightarrow \pi^*$	
5	[Cu(AZD2)Cl ₂]	<u>40000</u>	<u>250</u>	CT	1.95 S.P.
		<u>33333</u>	<u>300</u>	CT	
		<u>25000</u>	<u>400</u>	CT	
6	[Ni(AZD2)Cl ₂]	<u>45455</u>	<u>220</u>	CT	0 S.P.
		<u>25000</u>	<u>400</u>	${}^1A_{1g} \rightarrow {}^1A_{2g}$	
7	[Co(AZD2)Cl ₂]	<u>43478</u>	<u>230</u>	CT	3.89 T.h.
		<u>25000</u>	<u>400</u>	${}^4A_2(F) \rightarrow {}^4T_1(F)$	
8	[Mn(AZD2)Cl ₂]	<u>47619</u>	<u>210</u>	CT	
		<u>40000</u>	<u>250</u>	CT	
		<u>23809</u>	<u>420</u>	CT	

3.2 Infra-Red Spectroscopy

The infrared spectra assignment of the proposed structures of the Schiff base of azo-dye complexes was made through consideration of their infrared spectra. The IR spectra of the free ligand and metal complexes (Table 4) were carried out in the range 4000-400 cm^{-1} . The IR

spectrum of the all metal complex of (AZD1) ligand shows a sharp bands in range (1442 -1460) cm^{-1} , attributed to ν ($-\text{N}=\text{N}-$) stretching frequency of azo-dye⁽¹⁶⁾, which are shifted to higher frequency on going from the free ligand (at 1444 cm^{-1}). This is indicative of the coordination of the imine nitrogen to the metal. In the free ligands, the sharp bands observed in the range (3284) cm^{-1} is due to ν (OH) stretching frequency of carboxylic group⁽¹⁵⁾. In all complexes were moved to lower frequencies.

While IR spectra of the ligand (AZD2) exhibit band at (3219) cm^{-1} , which was assigned to the stretching vibration of (OH) phenol group⁽¹⁵⁾, while for all complexes are appear in range (3197-3205) cm^{-1} were shifted to lower frequency with change in shape in spectra of complexes, which is an indication of the engagement of this group in the coordination with the metal ion. The band in range (1587-1589) cm^{-1} for metal complexes are due to ν (C=N) stretching frequency of azomethin group, but for free ligand it appear at (1595) cm^{-1} shifted to lower frequency, which is an indication of the engagement of this group in the coordination with the metal ions. Metal oxygen, nitrogen and halogen bands further confirmed by the presence of the bands around (400-422), (447-469) and (526-586) cm^{-1} ^(16, 17) respectively.

Table (4): Selected IR bands of the azo- Schiff base ligands and its metal complexes (cm^{-1}).

No.	Compounds	ν (H-O)	ν (C=O)	ν (C=N)	ν (N=N)	ν (M-N)	ν (O-M)	ν (M-Cl)
	Ligand(AZD1)	3284	1697	1599	1425			
1	[Cu(AZD1)Cl ₂]	3275	-----	1603	1444	422	460	534
2	[Ni(AZD1)Cl ₂]	3305	1700	1593	1448	410	447	526
3	[Co(AZD1)Cl ₂]	3307	1705	1596	1442	407	453	532
4	[Mn(AZD1)Cl ₂]	3305	1700	1595	1460	407	459	560
	Ligand (AZD2)	3219	1734	1595	1441			
5	[Cu(AZD2)Cl ₂]	3199	1733	1585	1437	402	465	586
6	[Ni(AZD2)Cl ₂]	3203	1703	1587	1441	400	462	534
7	[Co(AZD2)Cl ₂]	3197	-----	1589	1441	402	469	530
8	[Mn(AZD2)Cl ₂]	3205	-----	1589	1437	402	461	526

3.4 Antibacterial activity:

In vitro antibacterial screening was performed by the agar disc diffusion method. The bacterial species used in the screening were gram-negative bacteria *Escherichia coli*, *Klebsiella*

pneumonia and gram-positive bacteria *Staphylococcus aureus* and *Staphylococcus haemolyticus*. The data are organized in table (5).

The free ligand (AZD1), (3) and (4) were sensitive against *Escherichia coli*. The free ligand (AZD1 and AZD2) (2), (5) and (8) were sensitive against *Staphylococcus aureus*. While the other complexes were resistant against *Klebsiella pneumonia* and *Staphylococcus haemolyticus*.

Table 5: Bacterial activity of ligands and its metal complexes

No.	Compound	<i>Escherichia coli</i>	<i>Klebsiella pneumonia</i>	<i>Staphylococcus aureus</i>	<i>Staphylococcus haemolyticus</i>
	Ligand(AZD1)	25 mm	R	20 mm	R
1	[Cu(AZD1)Cl ₂]	R	R	R	R
2	[Ni(AZD1)Cl ₂]	R	R	1 mm	R
3	[Co(AZD1)Cl ₂]	18 mm	R	R	R
4	[Mn(AZD1)Cl ₂]	27 mm	R	R	R
	Ligand (AZD2)	R	R	3 m	R
5	[Cu(AZD2)Cl ₂]	R	R	25 mm	R
6	[Ni(AZD2)Cl ₂]	R	R	R	R
7	[Co(AZD2)Cl ₂]	R	R	R	R
8	[Mn(AZD2)Cl ₂]	R	R	5 mm	R

Conclusion

This study has reported the synthesis and characterization of two Azo - Schiff base ligands containing (-N=N-) group, namely -{[2-Hydroxy-5-(4-nitro-phenylazo)-benzylidene]-amino}-benzoic acid (AZD1), 2-{[2-Hydroxy-5-(4-nitro-phenylazo)-benzylidene]-amino}-benzoic acid (AZD2). A novel ligands and its Cu(II), Ni(II), Co(II) and Mn(II) complexes were synthesized and characterized by various spectroscopic techniques. The results showed tetrahedral and square planer geometry around of the complexes with mole ratio (1:1) metal: ligand. According to Uv-visible spectroscopy, elemental analysis, IR spectroscopy, melting point, conductivity and magnetic susceptibility data the following structures are proposed for the synthesized metal complexes as shown in figure (1).

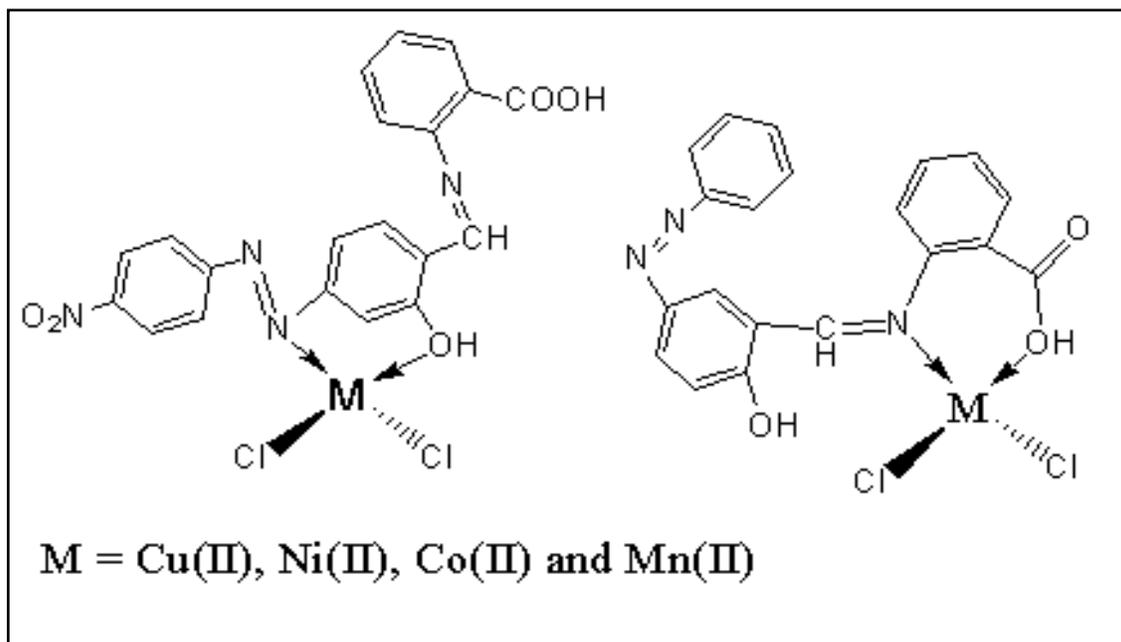


Figure (1): Proposed chemical structure for metal complexes.

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